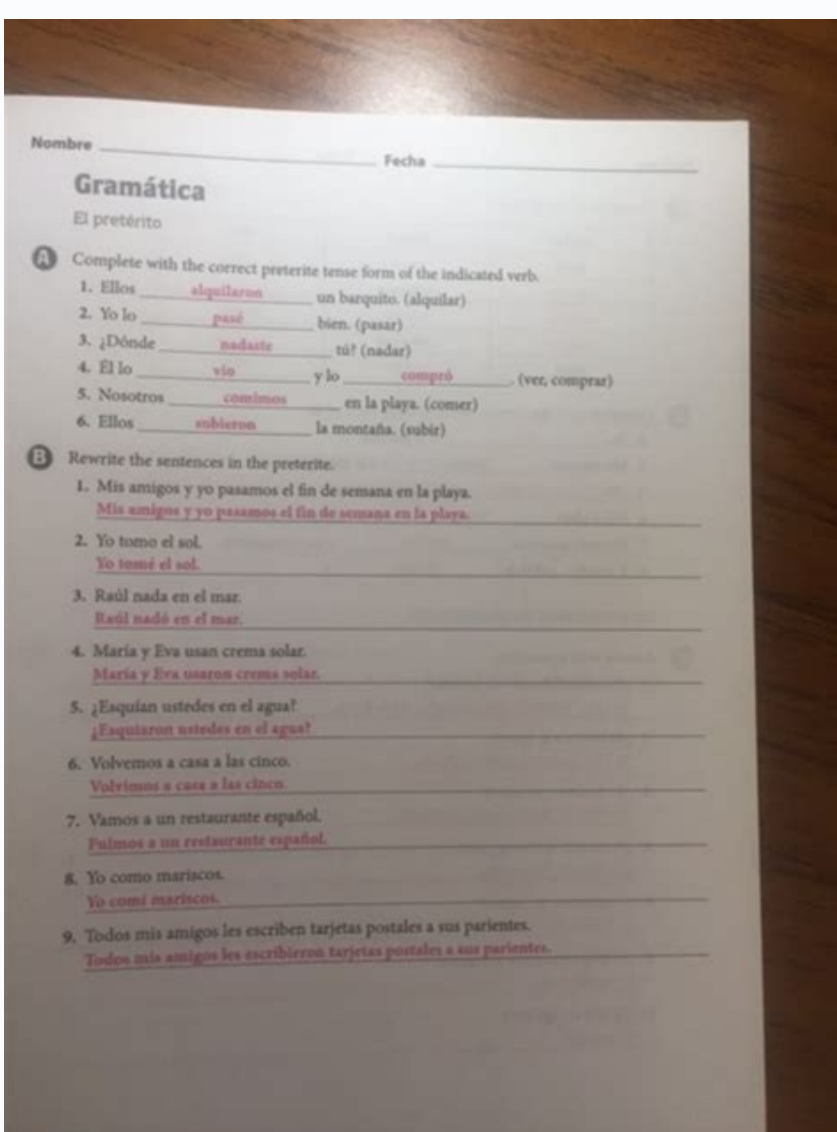
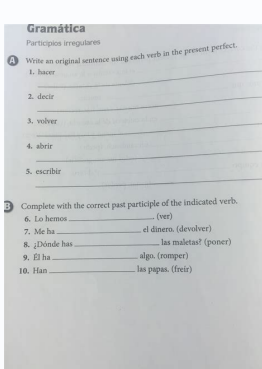




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Chemistry

(www.tiwariacademy.com)
(Chapter 2)(Solutions)
(Class 12)

Question 2.37:

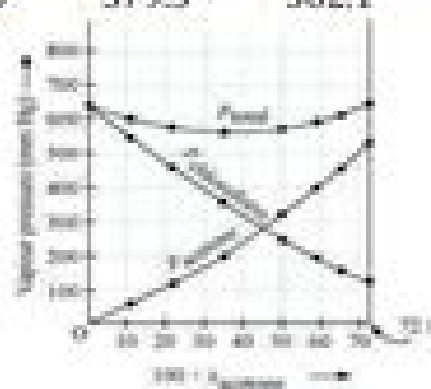
Vapour pressure of pure acetone and chloroform at 328 K are 741.8 mm Hg and 632.8 mm Hg respectively. Assuming that they form ideal solution over the entire range of composition, plot p_{total} , $p_{chloroform}$ and $p_{acetone}$ as a function of $x_{acetone}$. The experimental data observed for different compositions of mixture is.

$100 \times x_{acetone}$	0	11.8	23.4	36.0	50.8	58.2	64.5	72.1
$p_{acetone}$ /mm Hg	0	54.9	110.1	202.4	322.7	405.9	454.1	521.1
$p_{chloroform}$ /mm Hg	632.8	548.1	469.4	359.7	257.7	193.6	161.2	120.7

Answer 2.37:

From the question, we have the following data

$100 \times x_{acetone}$	0	11.8	23.4	36.0	50.8	58.2	64.5	72.1
$p_{acetone}$ /mm Hg	0	54.9	110.1	202.4	322.7	405.9	454.1	521.1
$p_{chloroform}$ /mm Hg	632.8	548.1	469.4	359.7	257.7	193.6	161.2	120.7
p_{total} (mm Hg)	632.8	603.0	579.5	562.1	580.4	599.5	615.3	641.8



It can be observed from the graph that the plot for the p_{total} of the solution curves downwards. Therefore, the solution shows negative deviation from the ideal behaviour.

Question 2.38:

Benzene and toluene form ideal solution over the entire range of composition. The vapour pressure of pure benzene and naphthalene at 300 K are 50.71 mm Hg and 32.06 mm Hg respectively. Calculate the mole fraction of benzene in vapour phase if 80 g of benzene is mixed with 100 g of toluene.

Answer 2.38:

Molar mass of benzene (C_6H_6) = $6 \times 12 + 6 \times 1 = 78 \text{ g mol}^{-1}$

Molar mass of toluene ($C_6H_5CH_3$) = $7 \times 12 + 8 \times 1 = 92 \text{ g mol}^{-1}$

Now, no. of moles present in 80 g of benzene = $\frac{80}{78} = 1.026 \text{ mol}$

And, no. of moles present in 100 g of toluene = $\frac{100}{92} = 1.087 \text{ mol}$

\therefore Mole fraction of benzene, $x_b = \frac{1.026}{1.026+1.087} = 0.486$

And, mole fraction of toluene, $x_t = 1 - 0.486 = 0.514$

It is given that vapour pressure of pure benzene, $p_b^0 = 50.71 \text{ mm Hg}$

And, vapour pressure of pure toluene, $p_t^0 = 32.06 \text{ mm Hg}$

Therefore, partial vapour pressure of benzene, $p_b = x_b \cdot p_b^0 = 0.486 \times 50.71 = 24.645 \text{ mm Hg}$

And, partial vapour pressure of toluene, $p_t = x_t \cdot p_t^0 = 0.514 \times 32.06 = 16.479 \text{ mm Hg}$

Hence, mole fraction of benzene in vapour phase is given by:

$$\frac{p_b}{p_b + p_t} = \frac{24.645}{24.645 + 16.479} = \frac{24.645}{41.124} = 0.599 = 0.6$$

Question 2.39:

The air is a mixture of a number of gases. The major components are oxygen and nitrogen with approximate proportion of 20% is to 79% by volume at 298

K. The water is in equilibrium with air at a pressure of 10 atm. At 298 K if the

Henry's law constants for oxygen and nitrogen are $3.30 \times 10^7 \text{ mm}$ and $6.51 \times 10^7 \text{ mm}$ respectively, calculate the composition of these gases in water.

Answer 2.39:

Percentage of oxygen (O_2) in air = 20 %

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